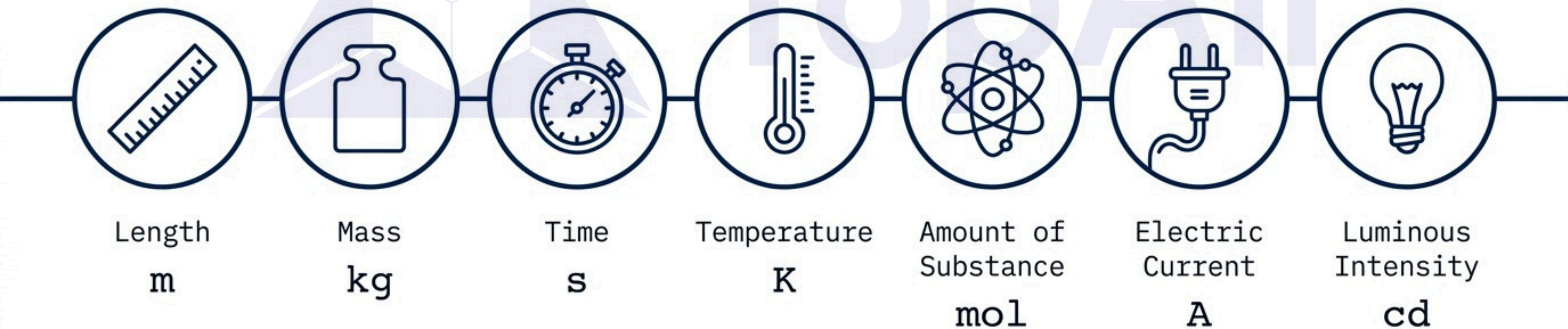


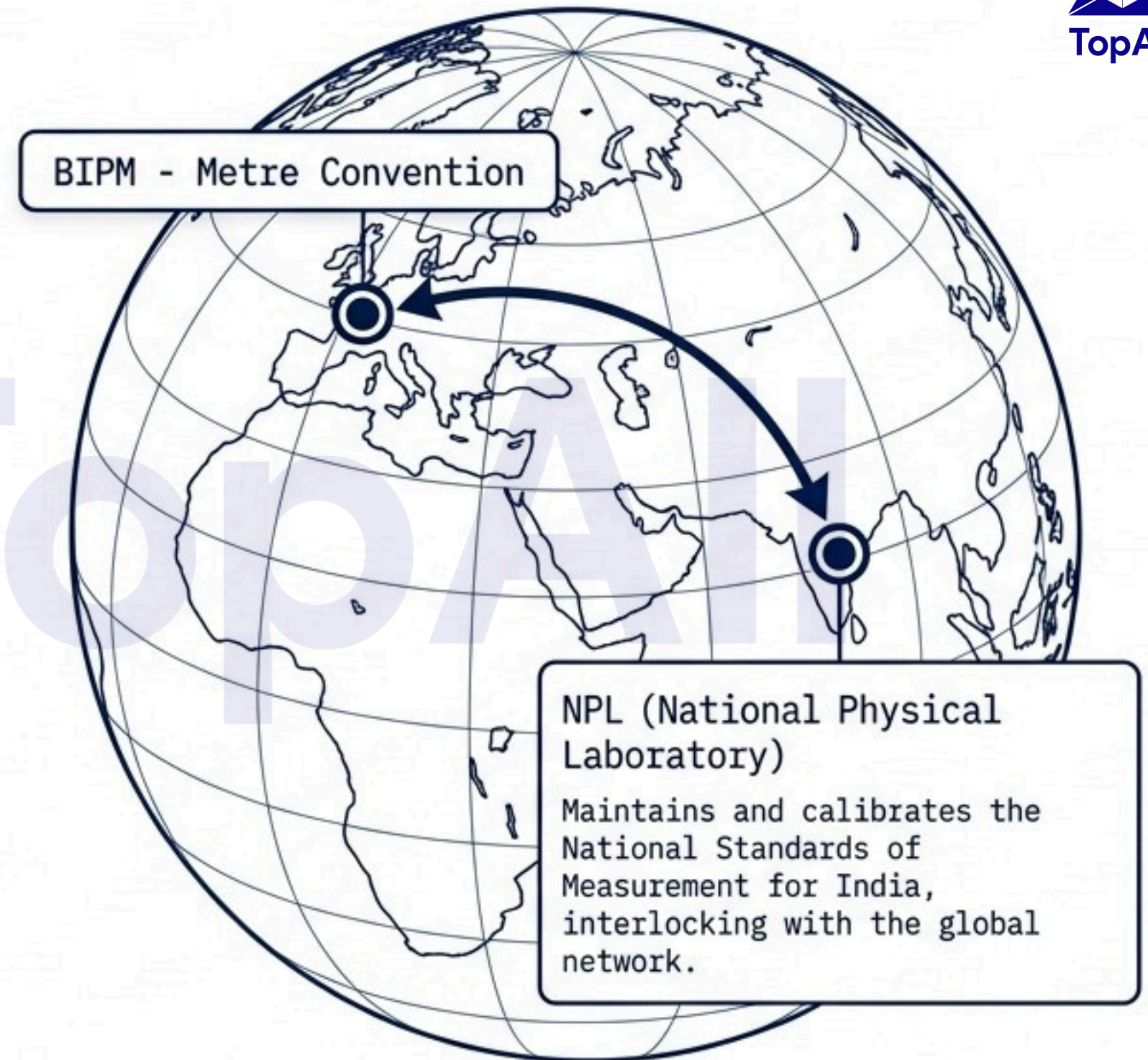
The Seven SI Base Units

The Fundamental Language of Chemistry & Physics



The Metre Convention (1875)

The International System of Units (SI) was established by diplomatic treaty in Paris. It replaced the chaos of localized, conflicting measurement systems with a single, elegant global standard for all scientific inquiry.



Physical Quantity	Symbol	SI Unit Name	SI Unit Symbol
Length	l	metre	m
Mass	m	kilogram	kg
Time	t	second	s
Electric current	I	ampere	A
Thermodynamic temp.	T	kelvin	K
Amount of substance	n	mole	mol
Luminous intensity	I _v	candela	cd



**You've Learned the Theory.
Let's Test Them Now.**

$$E=mc^2$$



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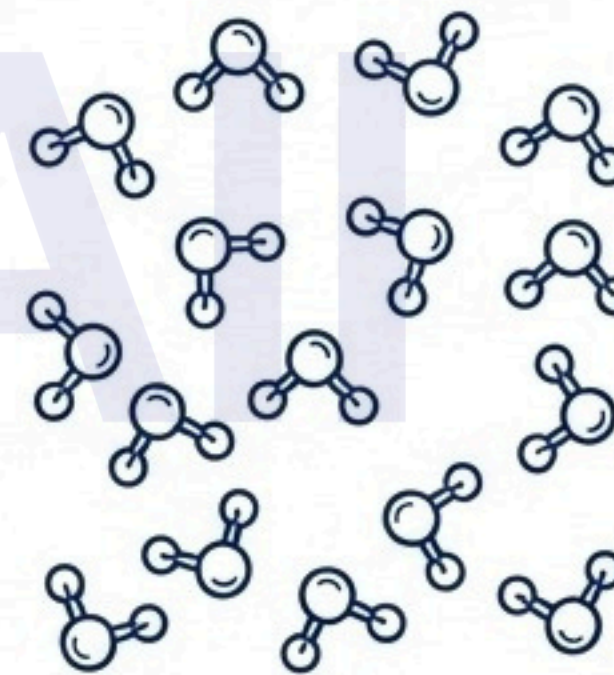
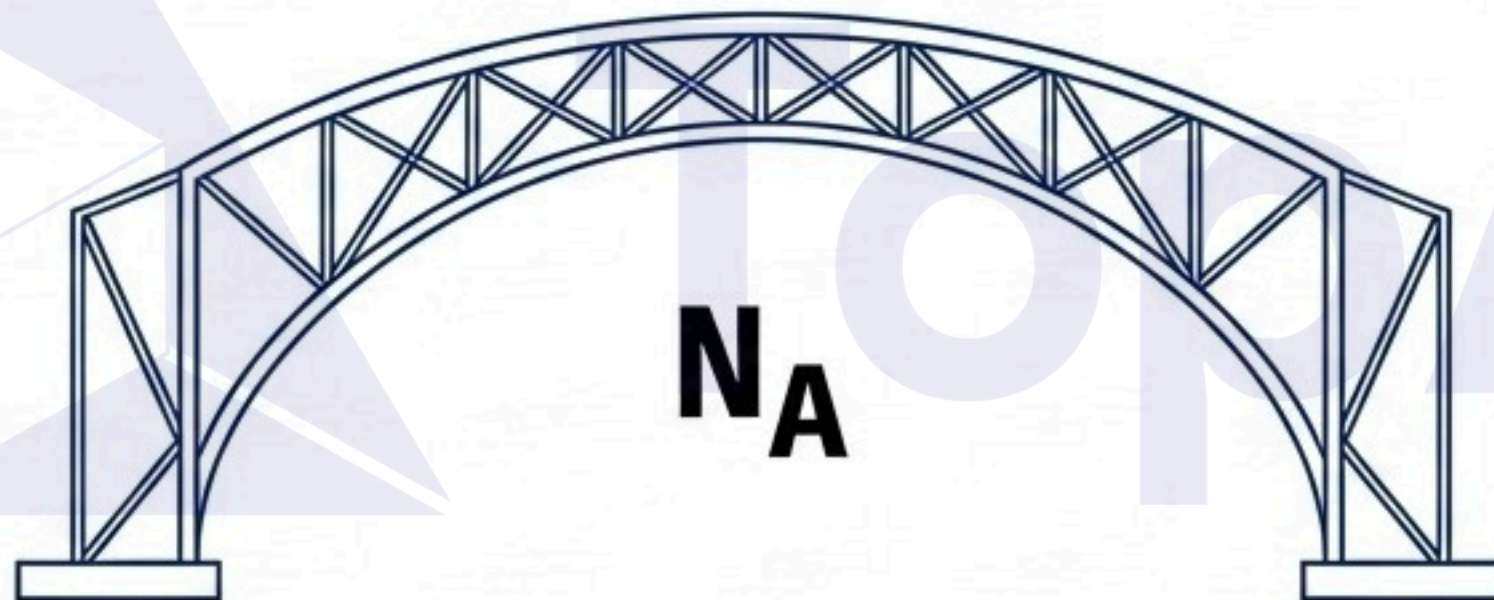
Key Highlights

- » Free Study Material, Videos & Infographics
- » Push Your Limits with Level-Based Challenges
- » Climb the Leaderboard Among Fellow Aspirants
- » Practice Unlimited MCQs for NEET & JEE
- » Get Clear Explanations, Not Just Marks
- » Track Your Performance Smarter



Amount of Substance (n)

$$1 \text{ mol} = 6.02214076 \times 10^{23} \text{ elementary entities}$$



Entities can refer to atoms, molecules, ions, or electrons. The mole guarantees the same number of entities regardless of the substance's identity.

JEE Trap Alert



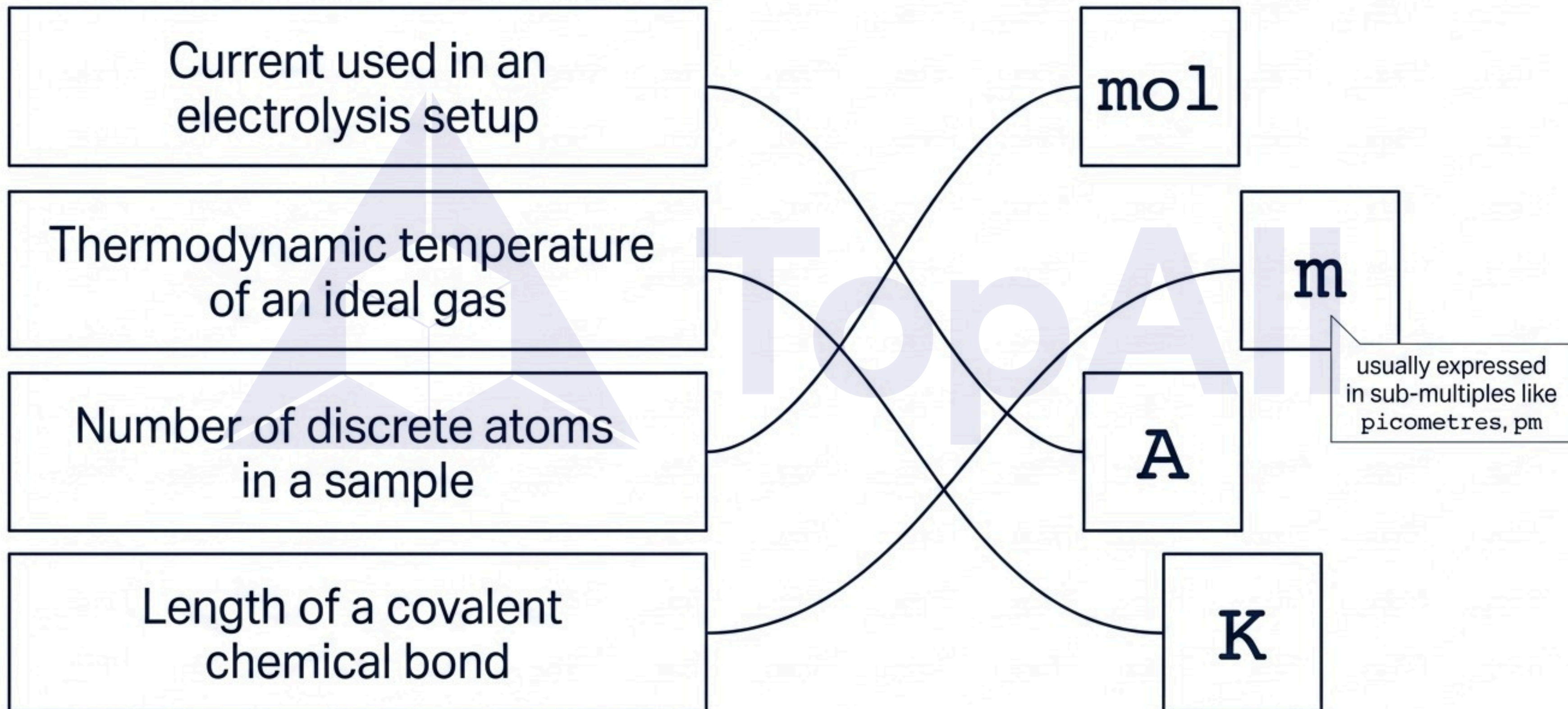
The Trap: Assuming gram is the SI base unit for mass because of the kilo prefix.

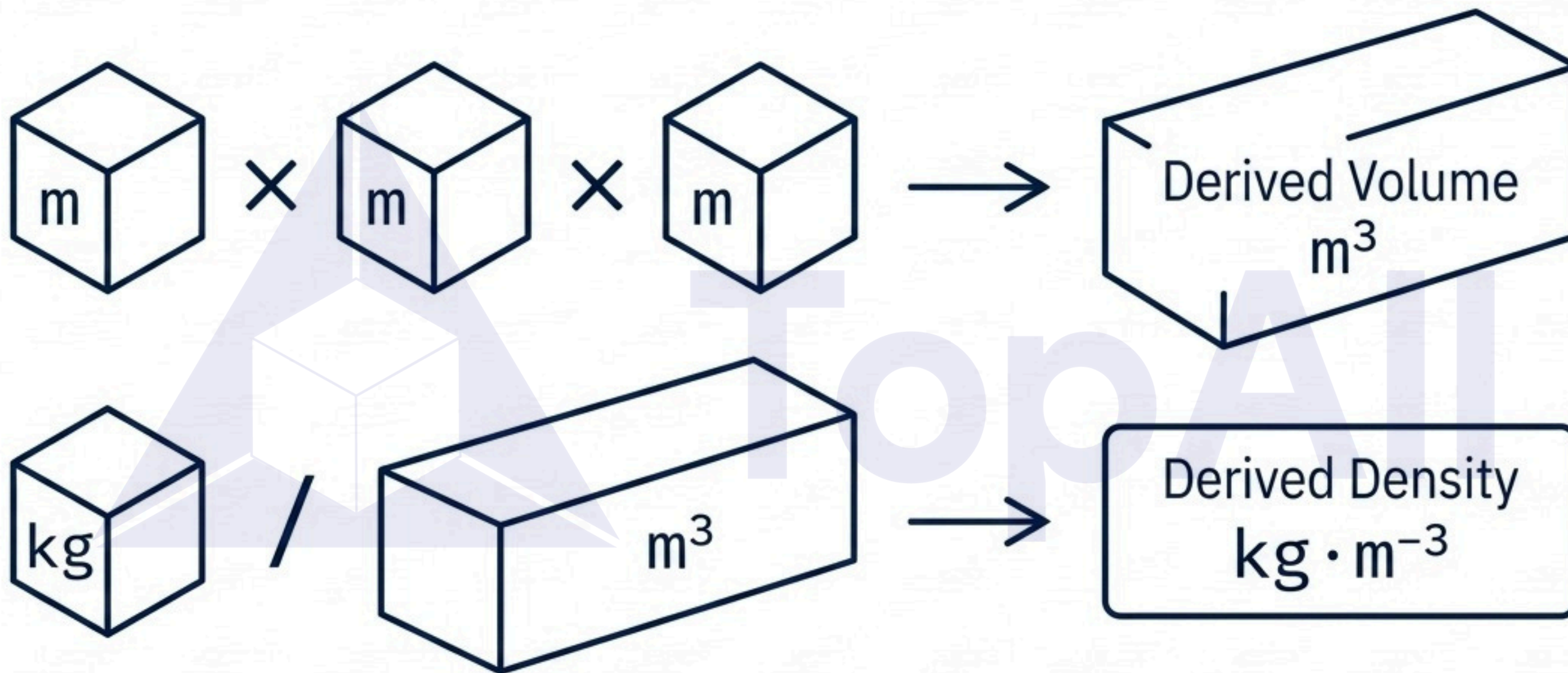
The Truth: The kilogram (kg) is the ONLY SI base unit that has a prefix built directly into its fundamental name.

Mass vs. Weight

Beware: Mass (constant amount of matter) $\neq$ Weight (gravity-dependent force). While chemistry labs use grams for practical small-scale convenience, stoichiometry in SI strictly grounds back to the kilogram.

Mini Application: Match the Quantity





All complex quantities in chemistry—from molarity to pressure—are simply **words** spelled out using the fundamental **alphabet** of the 7 SI base units.



You've Learned the Theory.

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