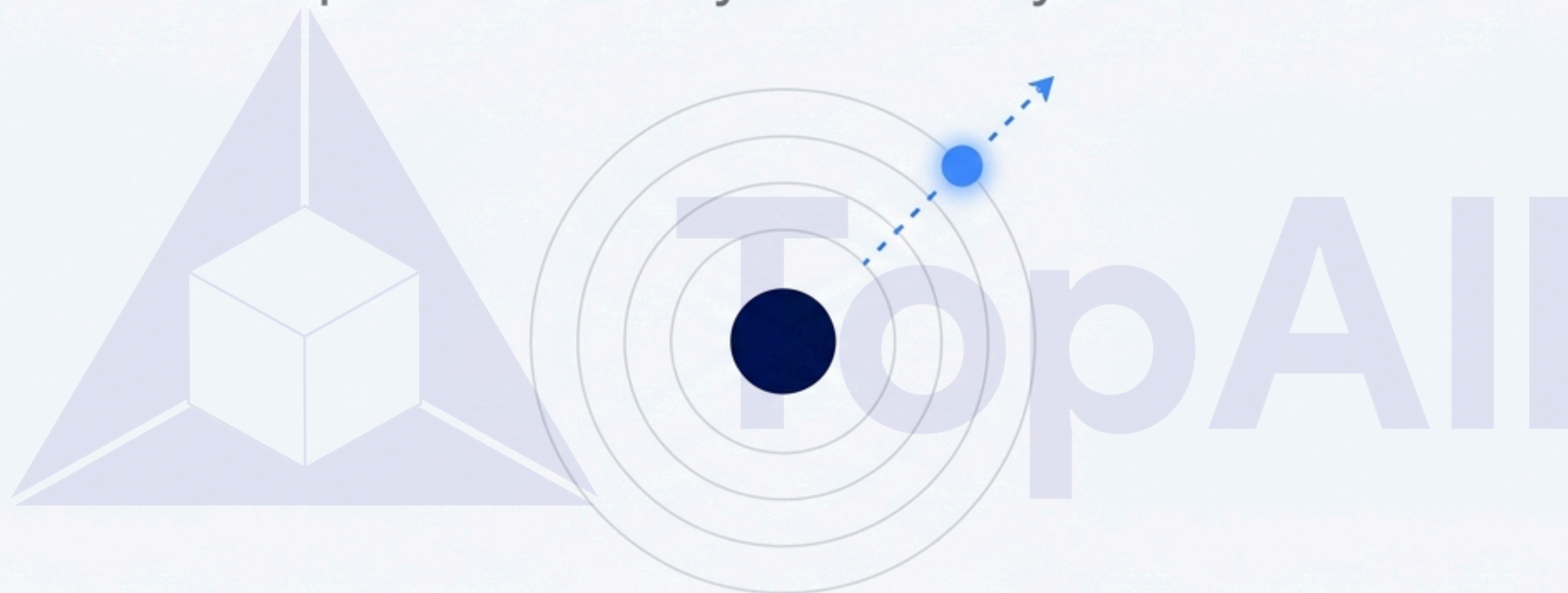


[Target Concept: s-Block]



Why s-Block Metals Lose Electrons Easily

Concept 8 • Periodicity & Reactivity Fundamentals



s-Block elements have strictly 1 or 2 valence electrons. Their chemistry is entirely driven by the energetic urge to shed them.



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$$E=mc^2$$



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Key Highlights

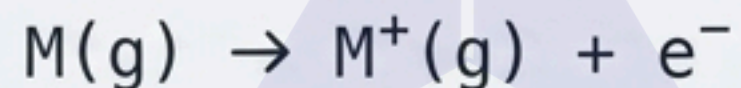
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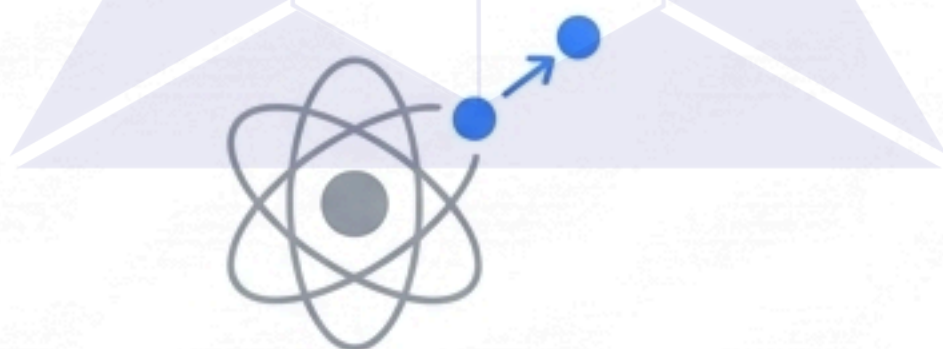
Ion Formation: The Drive for Stability

Shedding electrons to achieve Noble Gas configuration.

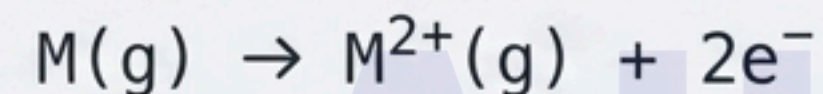
+1 Oxidation State



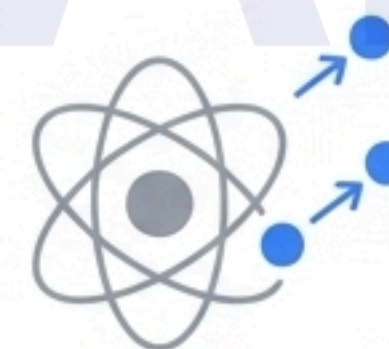
Achieves stable [Noble Gas] core.



+2 Oxidation State



Achieves stable [Noble Gas] core.



Result: s-Block elements almost exclusively form highly stable IONIC compounds.

Down-Group Trends: Size Dictates Reactivity

What happens as we move from top to bottom?



Radius & Shielding INCREASE

New principal energy levels are added. Inner core electrons shield the valence electron from the nucleus.

Ionization Enthalpy DECREASES

Increased distance and shielding mean less energy is required to remove the outermost electron.

Reactivity & Metallic Character INCREASE

Easier electron loss equals higher chemical reactivity.

The Exceptions: Lithium & Beryllium

First-element anomalies that break the rules.

EXAM TRAP: Do not apply standard s-block rules to Li and Be.

Tiny Size & High Charge/Radius Ratio

Extremely polarizing.

High Electronegativity

Form *covalent* compounds, not ionic.

No d-Orbitals

Maximum covalency is strictly limited to 4.

Key Concept: Diagonal Relationship
(Li matches Mg; Be matches Al)



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