

Defining Bond Order



In the Lewis description, Bond Order equals the total number of shared electron pairs between two atoms.

• 1 Shared Pair = Bond Order 1

Canonical Diatomic Examples

Hydrogen (H_2)



1 Shared Pair

Bond Order 1

Oxygen (O_2)



2 Shared Pairs

Bond Order 2

Nitrogen (N_2)

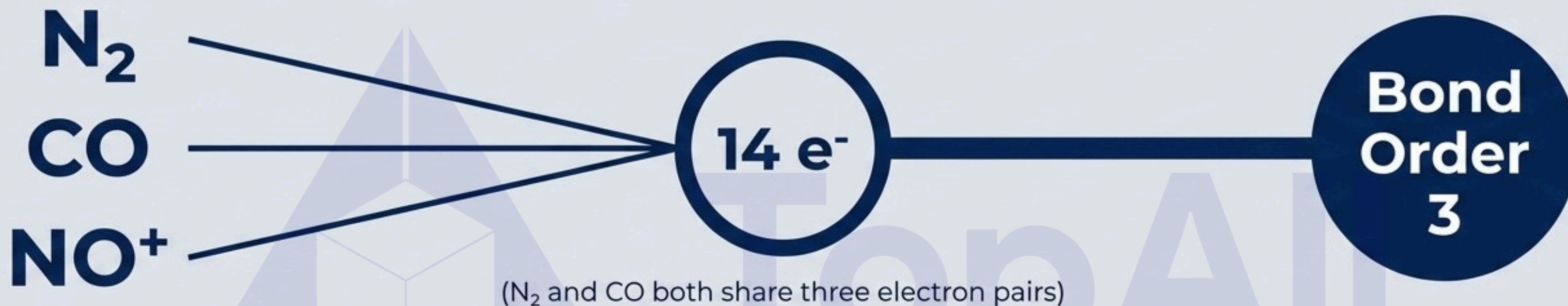


3 Shared Pairs

Bond Order 3

The Isoelectronic Rule

Molecules and ions with identical total electron counts share identical bond orders.





You've Learned the Theory. Let's Test Them Now.

$$E=mc^2$$



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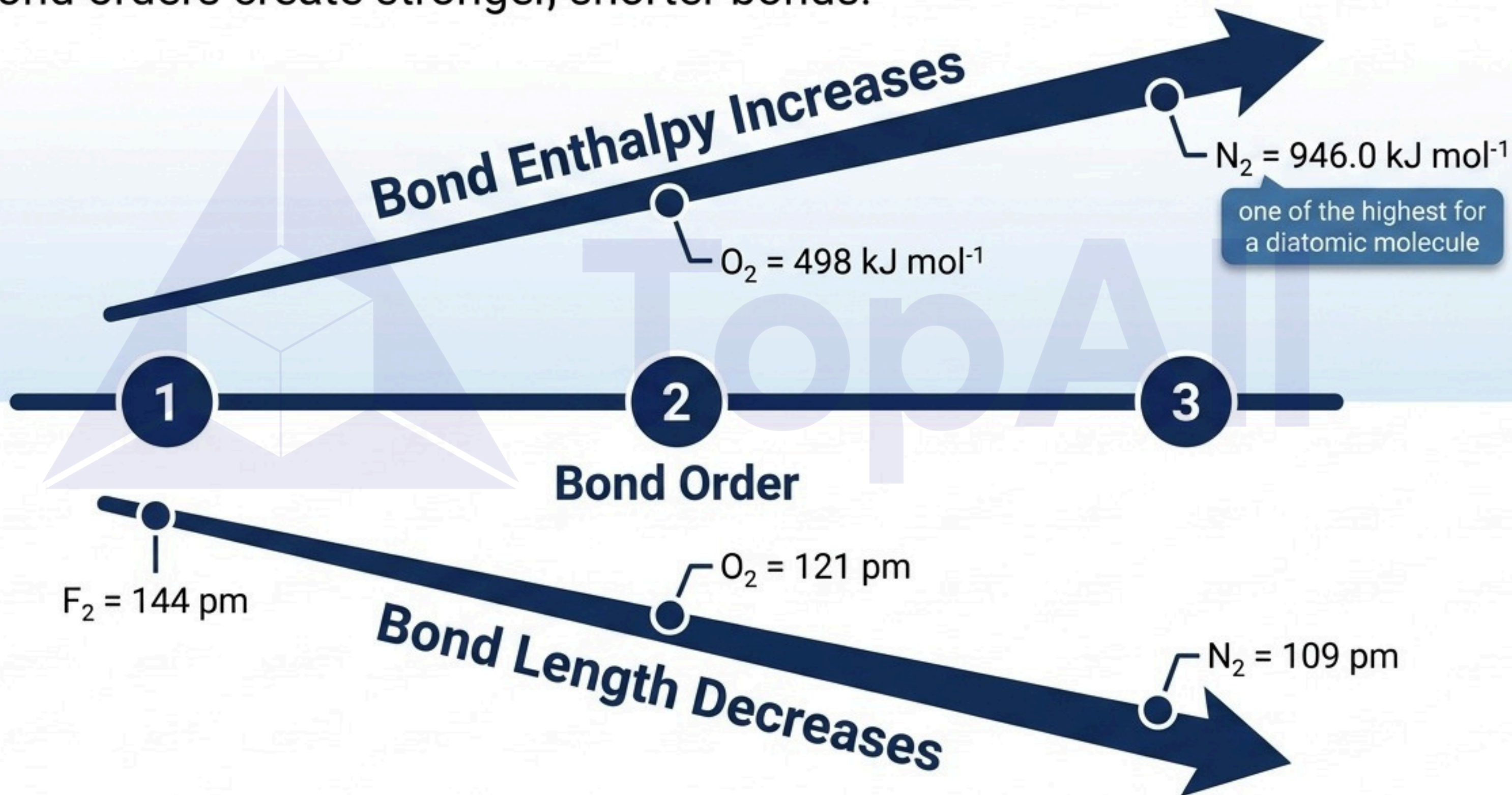
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Physical Properties Trend

Higher bond orders create stronger, shorter bonds.



Revision Frame: Bond Order

The Definition

Lewis Bond Order = Number of shared electron pairs.

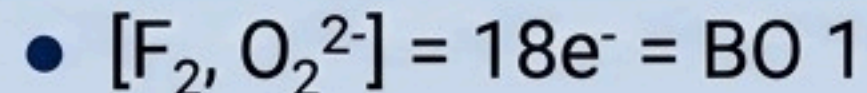
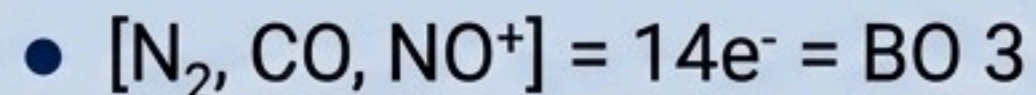


Core Examples



Isoelectronic Rule

Identical e⁻ count = Identical Bond Order.



Physical Properties



As **Bond Order** ↑, **Enthalpy** ↑ and **Length** ↓



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