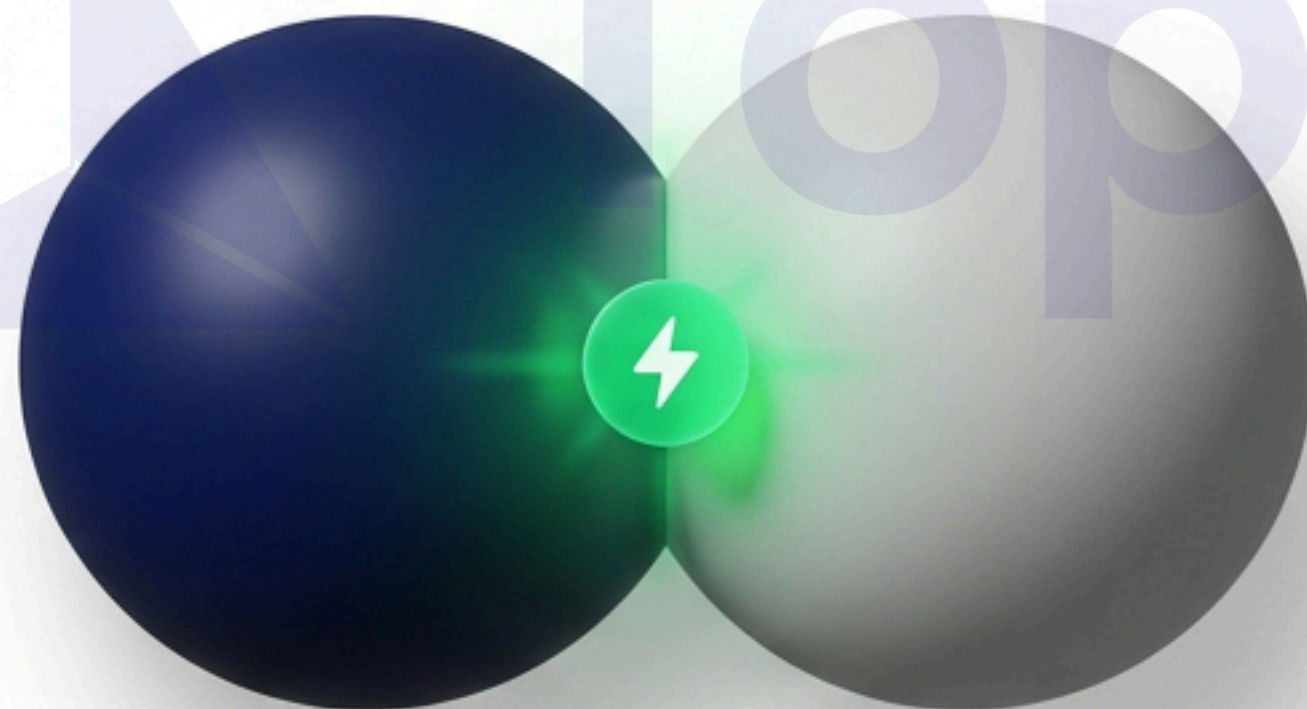


Combination Redox: When Elements Join

A redox reaction where two or more substances combine to form a single product, driven by the transfer of electrons.



The Core Formula

Reactant A

Reactant B

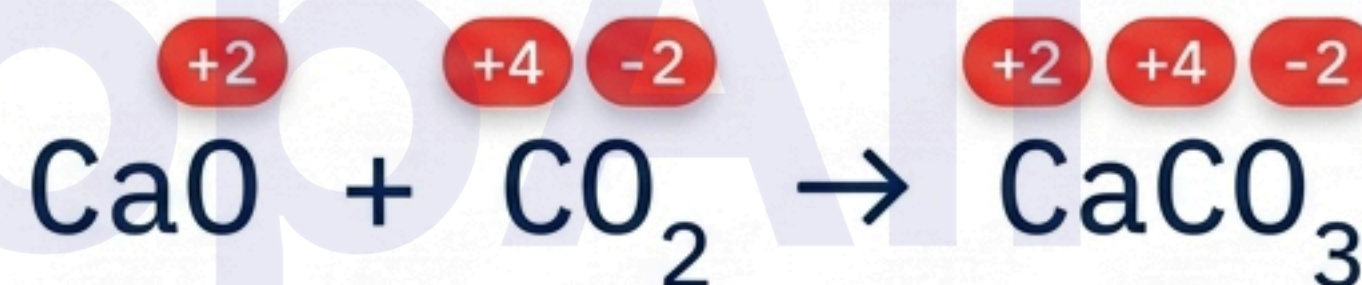
Product C



Two distinct entities fuse into a single unified compound.

The Golden Rule

For a combination reaction to be a redox reaction, at least one reactant must be in its elemental form.



No elemental reactant =
No electron transfer = NOT Redox.



**You've Learned the Theory.
Let's Test Them Now.**

$$E=mc^2$$



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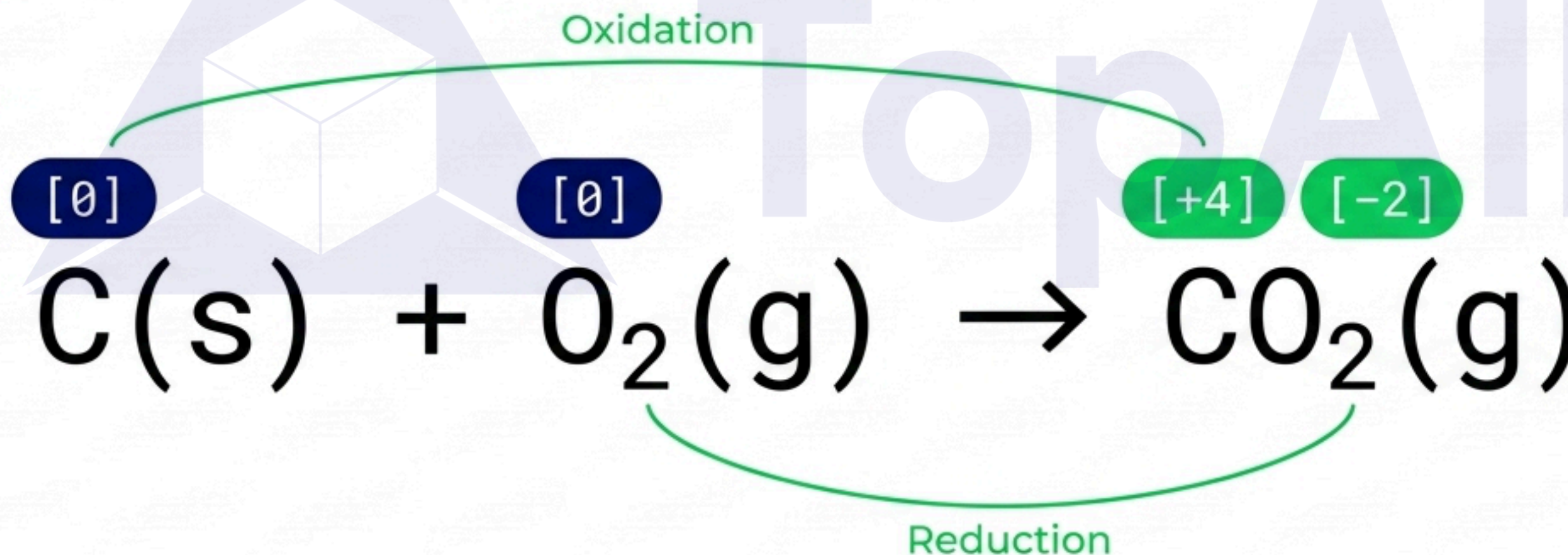
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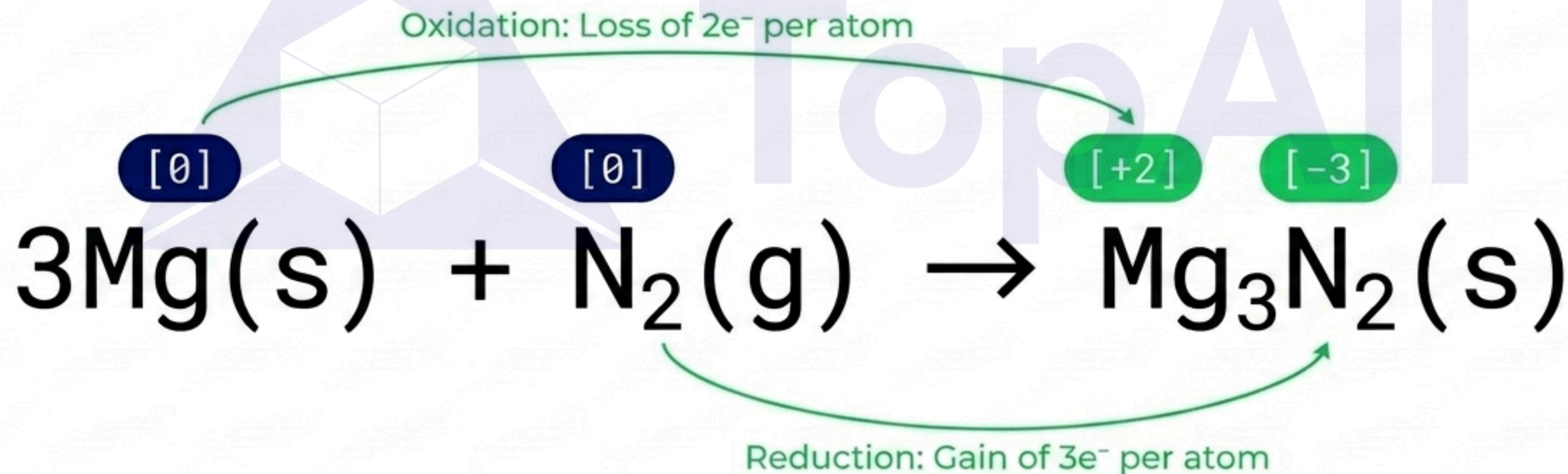
Application: Combustion

All combustion reactions utilizing elemental **dioxygen** are combination redox reactions.



Application: Synthesis

Combination redox also includes the direct synthesis of compounds using elements other than dioxygen.



Key Takeaways

The Form



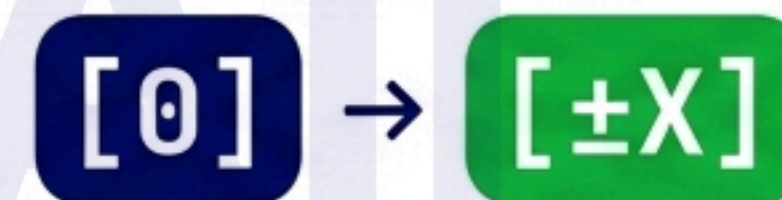
Multiple reactants combine to form one single product.

The Rule

Elemental Reactant

Either A, B, or both must be in an uncombined, elemental state.

The Shift



The oxidation state must shift from zero to a non-zero value, proving electron transfer.



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