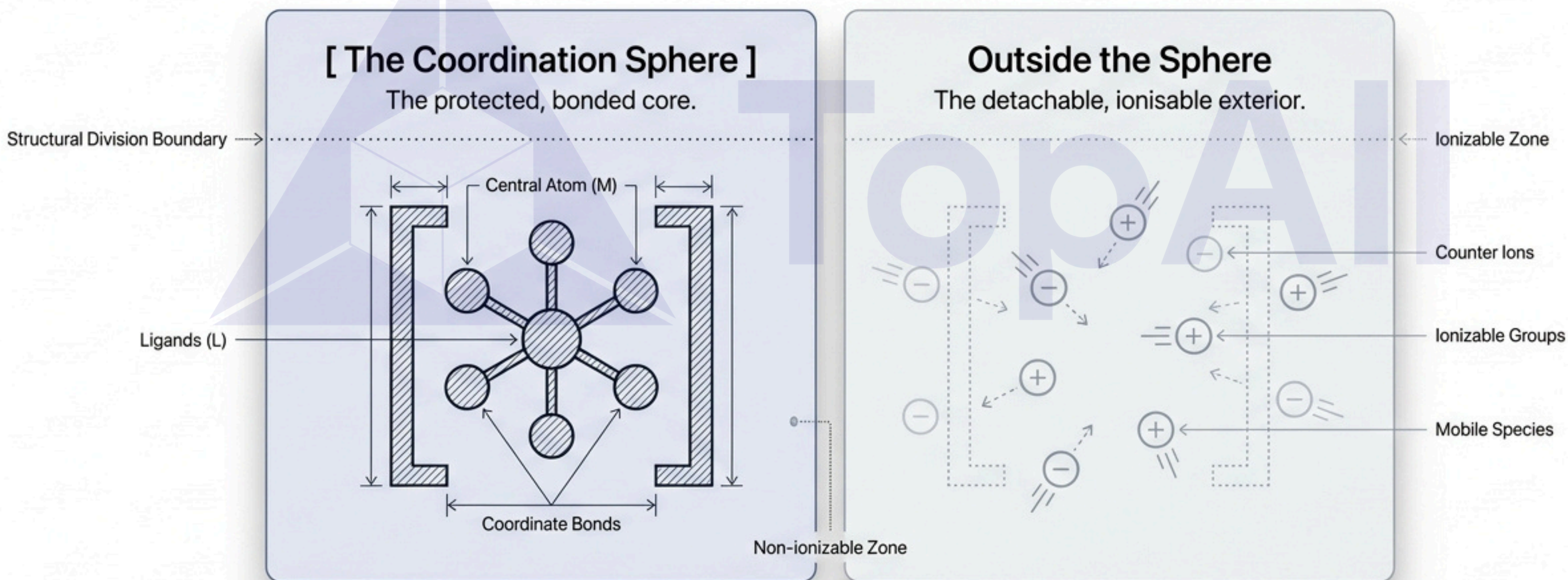


The Two Structural Zones of a Complex

Every coordination compound is divided into two distinct structural zones separated by a rigid boundary.



Decoding the Blueprint: $\text{K}_4[\text{Fe}(\text{CN})_6]$

The brackets define the absolute physical boundaries of the complex.



[COUNTER IONS]

K^+ ions sit strictly outside the brackets. They form the ionisable exterior and detach in solution.

[COORDINATION SPHERE]

The central Fe ion and its attached CN ligands are enclosed securely. This entire group exists collectively as $[\text{Fe}(\text{CN})_6]^{4-}$.

Coordination Sphere vs. Counter Ions

Coordination Sphere	Counter Ions
Enclosed securely in square brackets []	Written externally, outside the brackets
Central atom/ion + securely attached ligands	Mobile, ionisable groups
[Fe(CN)₆]⁴⁻	K⁺



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$$E=mc^2$$



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Master the Three Structural Principles

You are ready. Master these three principles to instantly identify components in any formula.



The Bracket Rule

If it is enclosed inside [], it is definitively defined as part of the coordination sphere.



The Ionisation Rule

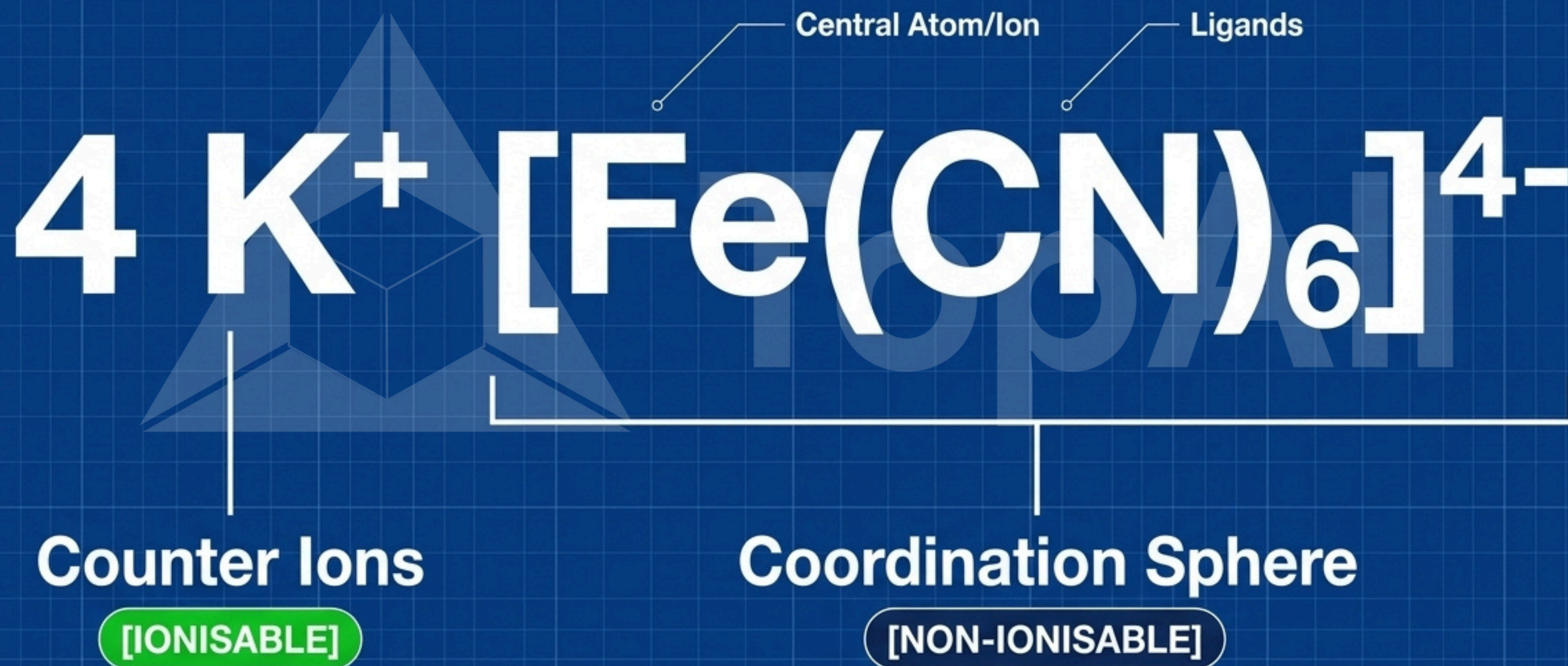
Only the groups written outside the brackets (the counter ions) are ionisable in aqueous solution.



The Collective Identity

The central metal and its attached ligands act permanently as a single, collective structural entity.

The Anatomy of Potassium Hexacyanoferrate(II)





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