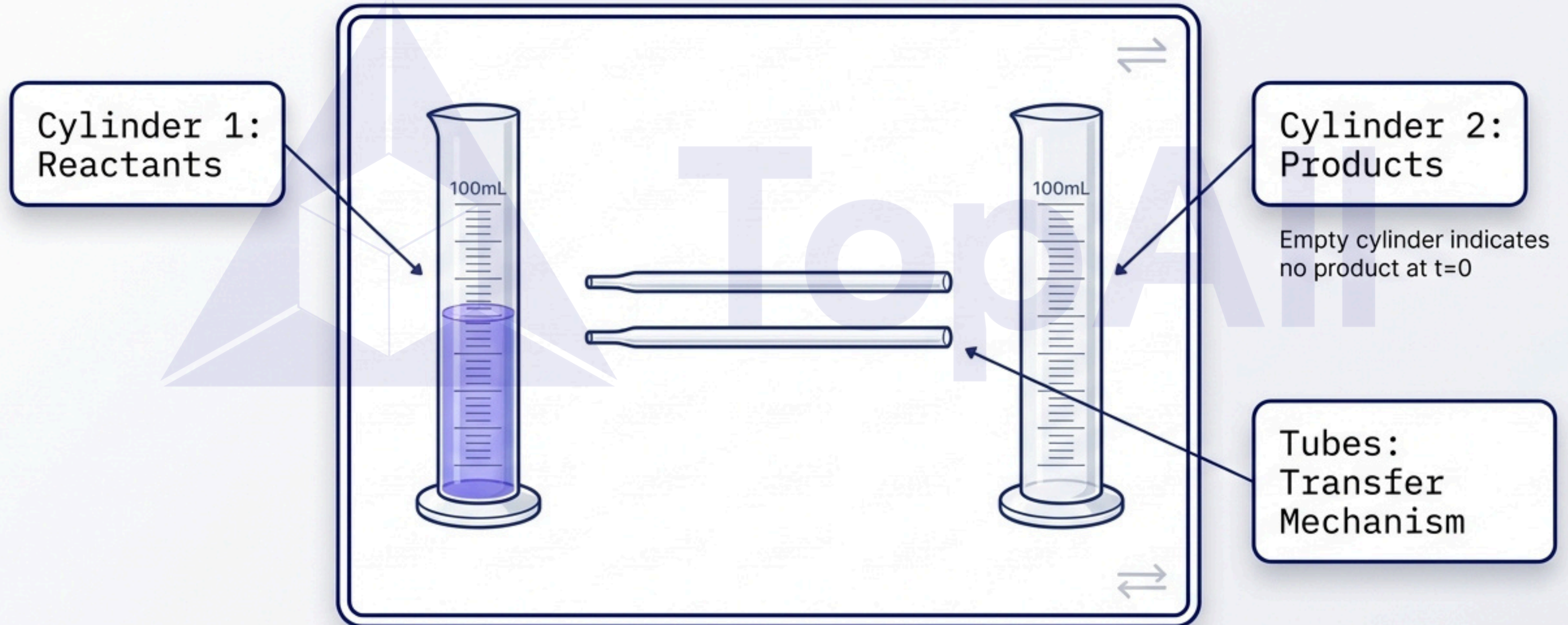


Dynamic Equilibrium Activity

A visual demonstration of reversibility using the coloured water experiment.



Preparing the Equilibrium System





**You've Learned the Theory.
Let's Test Them Now.**

$$E=mc^2$$



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Practice Now



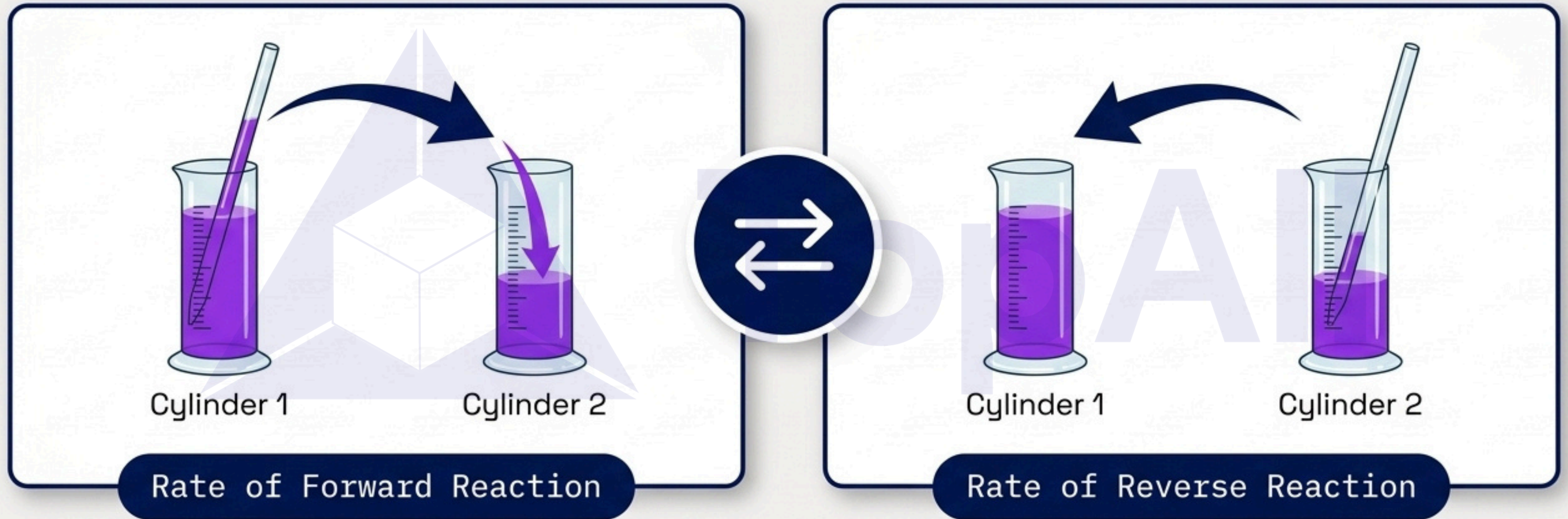
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Key Highlights

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- » Get Clear Explanations, Not Just Marks
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The Continuous Exchange



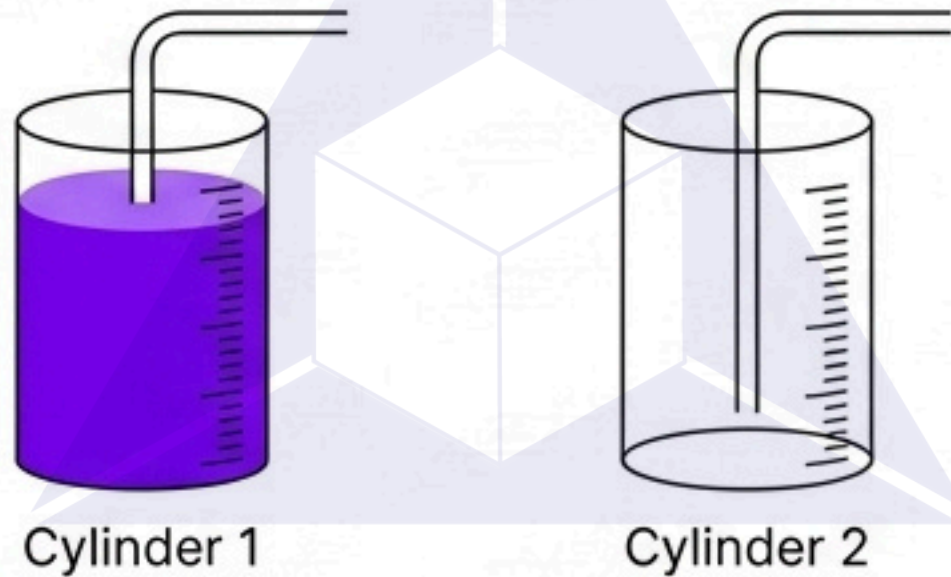
Key Takeaway Panel: The process of transfer continues simultaneously in both directions. One tube moves water left-to-right, while the other moves it right-to-left.

THE OBSERVATION

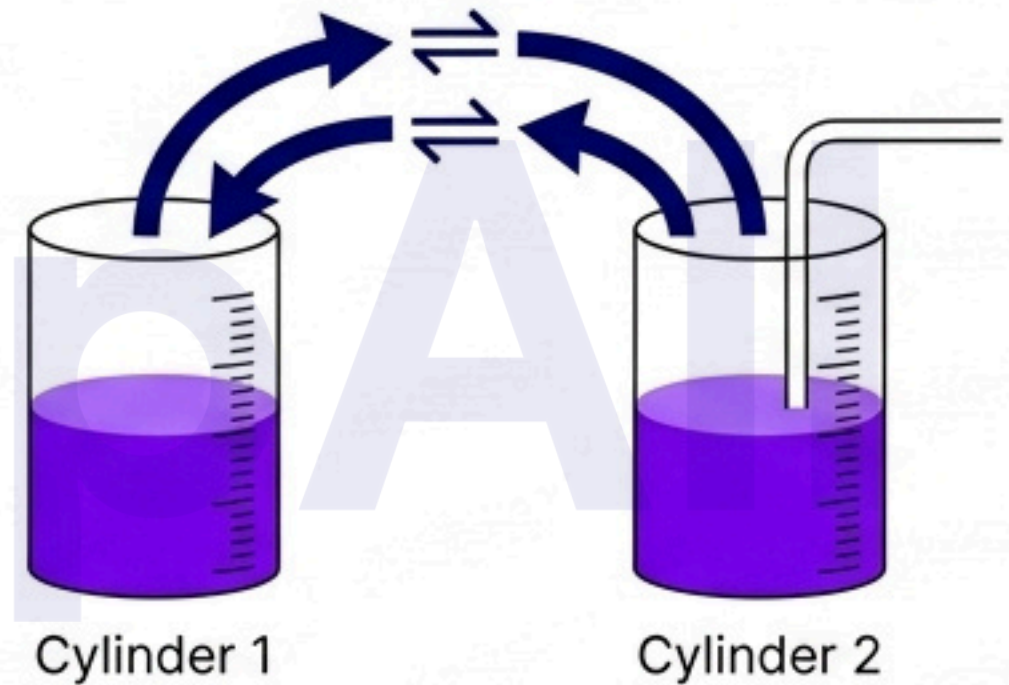
Macroscopic Constancy



Initial Stage



Equilibrium Attained



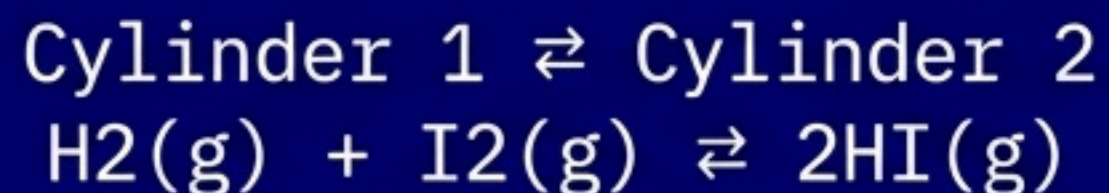
The levels of coloured water stop changing. However, the physical transfer of water is still occurring in both directions.

Not Static Rest, But Dynamic Balance

Physical elements		Chemical theory
Water Level	$\leftrightarrow$	Concentration of Reactants/Products
Empty Cylinder 2	$\leftrightarrow$	Zero Products at initial stage ($t=0$)
Tube Transfer	$\leftrightarrow$	Reaction Rate
Constant Levels	$\leftrightarrow$	Equilibrium State ($\text{Rate}_{\text{fwd}} = \text{Rate}_{\text{rev}}$)

Key Takeaways Panel

Static vs. Dynamic: The system is NOT at rest. The forward and backward processes continue at the exact same rate, resulting in no net change in concentration.





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